$\cancel{\Delta H_{fus} = MC\Delta T}$

$\cancel{12.327 = x \times 11.2 \times 350}$

a) $$\ln\left(\frac{P_2}{P_1}\right) = \frac{\Delta H_{vap}}{R}\left(\frac{1}{T_2} - \frac{1}{T_1}\right)$$

$$= \underline{29.1\ kJ/mole}$$

$$\ln\frac{P_2}{P_1} = \frac{29.1\ kJ/mole}{58.0791}\left(\frac{1}{178.7} - \frac{1}{296}\right)$$

$$= 0.50104\ (0.005596 - 0.003378)$$

$$= 0.50104 \times 0.002218$$

$$\ln\frac{P_2}{P_1} = 0.0011131$$

$$P = \underline{\underline{1.0011\ atm}}$$

Generated by CamScanner